

**YEAR: 9**

**SUBJECT: SCIENCE**

**TEST: Chemical reactions**

**TIME: 45 mins**

**QUESTIONS: Part A: Multiple Choice Questions (5 marks)**

**Part B: Short Answer Questions (28 marks)**

**TOTAL MARKS: 33 marks**

**DO NOT WRITE ON OR MARK THIS PAPER**

**SECTION ONE: Multiple Choice Questions (1 mark each)**

**Answer this section on the separate multiple – choice answer sheet**

1. Copper reacts with sulfur dioxide to form copper sulfide and oxygen gas. The reactants for this reaction are:

A) copper

B) copper and sulfur dioxide

C) copper sulfide and oxygen gas

D) copper, sulphur dioxide, copper sulfide and oxygen gas

1. Which of the following is the balanced chemical equation showing the corrosion of aluminium?

A) Al + O2 → Al2O3

B) 4Al + O2 → 2Al2O3

C) Al + 3O2 → 2Al2O3

D) 4Al + 3O2 → 2Al2O3

1. When petrol explodes, it releases energy in the form of heat and light. This reaction is an example of:

A) an endothermic reaction

B) an exothermic reaction

C) a neutralisation reaction

D) a corrosion reaction

1. Which of the following reactions shows the rusting of iron?

A) 2Cu + H2O + CO2 + O2 → Cu(OH)2 + CuCO3

B+) 2Ag + H2S → Ag2S + H2

C) 2Na + 2H2O → 2NaOH + H2

D) 4Fe + 3O2 + 2H2O → 2Fe2O3.H2O

1. An acid reacts with a metal. Identify its products.

A) a salt + hydrogen gas

B) a salt + water

C) a salt + water + carbon dioxide

1. D) glucose + oxygen gas



**SEMESTER Two 2017**

**Chemical reactions:**

**ANSWER BOOKLET**

**NAME:**

**FORM:** **DATE:**

Multiple Choice Short Answer Total

**/33**

**/28**

**/5**

**SECTION ONE:** Multiple choice answers

Cross (X) through the correct answer.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| 1 | a | b | c | d |
| 2 | a | b | c | d |
| 3 | a | b | c | d |
| 4 | a | b | c | d |
| 5 | a | b | c | d |

**SECTION TWO: Short Answer (28 marks)**

Answer the questions in the spaces provided.

**Question 5 (4 marks)**

Ethane (C2H6) reacts with oxygen gas (O2) to form carbon dioxide (CO­2) and water vapour (H2O).

**a)** **Identify** the reactants of this reaction.

**b)** **Identify** the products of this reaction.

**c)** **Construct** a word equation for this reaction.

**d)** **Construct** an unbalanced formula equation for this reaction.

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**Question 6 (5 marks)**

Magnesium burns in oxygen gas to form magnesium oxide. Its unbalanced formula equation is:

Mg + O2 → MgO

1. **Use** this equation to **identify** the chemical formula for magnesium oxide.
2. **Construct** a word equation describing this reaction.
3. **Identify** which element (Mg or O) is unbalanced in the above equation.
4. **Balance** the above equation.
5. **Identify** the type of reaction that has taken place.

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**Question 7 (4marks)**

**Classify** each of the following reactions as combustion, corrosion, acid/metal, neutralisation, or acid/carbonate reactions:

2HCl + Mg → MgCl2 + H2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

C6H12O6 + 6O2 → 6CO2 + 6H2O ­­­­­­­­­­­­­­\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

2HCl + CaCO3 → CaCl2 + H2O + CO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

H2SO4 + 2NaOH → Na2SO4 +2H2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Question 8 (4 marks)**

Two things can happen to an iron/steel shipwreck when it settles on the ocean floor. It can stay on the floor OR it can sink into the mud. If it stays on the ocean floor, it will rust. If covered by mud, it is sometimes preserved without any rusting.

The rusting of iron is shown in the equation:

4Fe + 3O2 + 2H2O → 2Fe2O3.H2O

**a)** **Use** this equation to **identify** the chemical formula for rust.

**b)** **List** the three reactants needed for rust to form.

**c)** **Propose** a reason why iron might not rust if the shipwreck is covered by mud.

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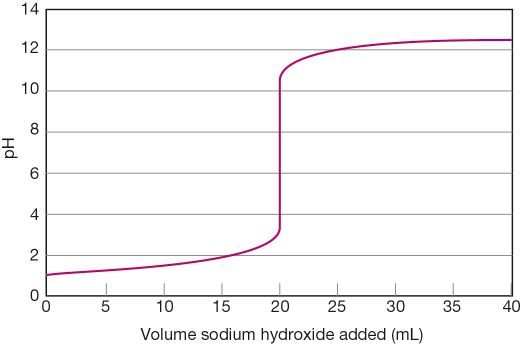
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**Question 9 (5 marks)**

Use this graph to answer the following questions:

Increasing volumes of sodium hydroxide (NaOH) were added to a solution of hydrochloric acid (HCl). As it was added, the pH changed according to the graph shown below.



1. State the pH of the solution at the start of the experiment.
2. Determine the volume of sodium hydroxide that is needed to be added to exactly neutralise the acid.
3. Determine the pH when 16 mL of sodium hydroxide was added.
4. Determine the total volume of sodium hydroxide added in the experiment.
5. At the end of the above experiment, what type of solution was formed?

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**Question 10**

Acids are commonly found around us. Bases can be described as the ‘chemical opposite’ of acids.

1. List the main properties of **acids** and **bases**. **(4 marks)**
2. What is the difference between a strong and a weak acid? (**2 marks)**

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Please check your work / complete any unanswered questions.